

ALUMINIUM HALIDE-THIOL SYSTEM: A USEFUL REAGENT FOR DEMETHYLATION OF ALIPHATIC AND AROMATIC METHYL ETHERS AND DEMETHYLENATION OF METHYLENEDIOXY COMPOUNDS

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A new efficient procedure for demethylation of methyl ethers and demethylenation of methylenedioxy compounds under mild conditions has been exploited by the use of aluminium halide-thiol system.

We have published a reagent system, BF_3 -etherate and thiol, for demethylation of aliphatic methyl ethers¹ and its application.² This reagent system has also been shown to be effective for debenzylolation.³ Now, we wish to report an exploitation of a new reagent system, AlX_3 and EtSH. As shown in Table 1, demethylation with aluminium halide-thiol system was tried with success on several aliphatic and aromatic methyl ethers. Also for the aromatic methyl ethers, demethylation proceeded very easily. (*Cf.* Selective demethylation of 17-OMe group in estradiol dimethyl ether (4) with BF_3 -etherate-thiol¹). A partial demethylation was successful on treatment of dimethyl ether (7) with AlCl_3 at -15°C for 5 min, which was applied to the total synthesis of lythranidine.⁴ Selective demethylation of the aromatic methyl ether from the compound having both aromatic methoxy and aliphatic or aromatic alkoxy carbonyl groups was also successful under the conditions indicated [see reactions of (8), (9), and (10)], although this system is available for dealkylation of esters under stronger conditions.⁵ By contrast, demethylation of the methoxycarbonyl group has been known to be prior to demethylation of methyl ether in (8), when RS^- was used.⁶

Application of this system was expanded to demethylenation of the methylenedioxy group. The methylenedioxy compounds, (21), (22), and (23), on treatment with AlBr_3 -EtSH gave satisfactory yields of the corresponding catechols in a short time under mild conditions, as summarized in Table 2.

These reactions proceed *via* initial coordination of "hard" Lewis acid [Al(III)] to "hard" base (oxygen) followed by attack of thiol, a "soft" base, to the activated less hindered carbon

atom ("soft" acid) through an S_N2 -type pattern.

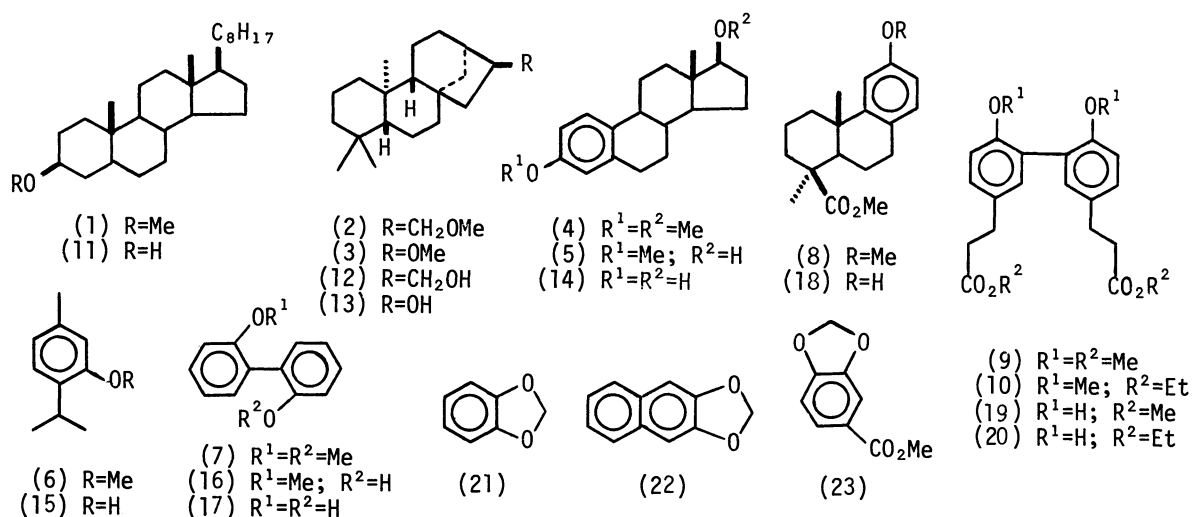
Table 1. Demethylation of aliphatic and aromatic methyl ethers with AlX_3 in EtSH

Substrate	AlX_3 (mol equiv.)	Temp.	Reaction time (hr)	Product and yield (%)
(1)	$AlBr_3$ (1.8)	r.t.	3	(11) 98.4
(2)	$AlBr_3$ (3)	r.t.	14 ^a	(12) 98.3
(3)	$AlBr_3$ (3)	r.t.	< 1	(13) 90.0
(4)	$AlBr_3$ (5)	r.t.	< 1	(14) 94.3
(5)	$AlCl_3$ (3.4)	r.t.	< 0.5	(14) 97.5
(6)	$AlCl_3$ (2)	0°	< 0.5	(15) 95.0
(7)	$AlCl_3$ (3) ^b	-15°	0.1	(16) 80, (17) 5.0
"	$AlCl_3$ (2.7)	r.t.	< 0.5	(17) 97.4
(8)	$AlCl_3$ (3) ^b	0°	6	(18) 98.6
(9)	$AlCl_3$ (5) ^b	0°-r.t.	2.5	(19) 98.0
(10)	$AlCl_3$ (5) ^b	r.t.	2	(20) 95.2

- a Because of water of crystallisation in the substrate, the reaction took such a long time.
 b Ethanethiol (5-10%) in dichloromethane was used.

Table 2. Demethylenation of methylenedioxy compounds with $AlBr_3$ in EtSH

Substrate	Mol equiv. of $AlBr_3$	Temp.	Reaction time (hr)	Yield (%) of catechol
(21)	2.4	0°	< 0.5	78.4
(22)	2.4	0°	< 0.5	73.0
"	1.2	r.t.	2	60.2
(23)	4	0°	1	72.4



References and notes

- 1 M. Node, H. Hori, and E. Fujita, *J. C. S. Perkin I*, 1976, 2237.
- 2 E. Fujita, M. Node, and H. Hori, *J. C. S. Perkin I*, 1977, 611.
- 3 K. Fuji, K. Ichikawa, M. Node, and E. Fujita, unpublished.
- 4 K. Fuji, K. Ichikawa, and E. Fujita, unpublished.
- 5 It will be published elsewhere.
- 6 T. R. Kelly, H. M. Dali, and W. -G. Tsang, *Tetrahedron Letters*, 1977 3859.

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